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CHM191
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2. Of all the oxidizable species available for use, iodine titration are usually used with the___

sulphate ions

--->> Thiopsulphate

hydrogensulphate

thioesters

3. The oxidation potentials of Na is_____

11

23

271

--->> 2.71

4. The starch solution forms a _____complex with the tri-iodide ion during the titration that is rapidly discharge at the end point

Light blue

Blue

--->> Blue black

Brick red

5. For this reason and in order to avoid complications in the end point detection, the starch indicator is added at the_____ colouration of the solution which is near the end point

Yellow

--->> Light yellow

Colourless

Dark yellow

6. The most commonly used external indicator for iodine titration is_____

Glucose solution

hydrogen chloride

Protein solution

--->> starch solution

7. Complete the colour change in iodometry analysis.
_____ (initially) $\hat{+}$ '

Red-orange

Dark-brown

Brown

--->> Red-brown

8. In the _____, standard solutions of iodine are used to estimate directly the concentrations of some oxidizable species.

--->> Direct method

indirect method

standard method

calibrated method

9. The _____ is the amount of heat released or observed for a given amount of reactants or products

--->> heat of reaction

heat of combustion

heat of neutralization

heat of formation

10. Thus using starch as indicator, the colour variations of the solution depend on _____ formation of the starch with iodide ion formed in the last stage of the titration.

--->> Complex

Colloid

Hydrolysed

Navy blue